ADVANCED ANSWER KEY

2021

CHEMISTRY Paper-2 QUESTION WITH SOLUTION

32700+ SELECTIONS



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(VJ Sir)

Exp. : 18 yrs



Exp. : 16 yrs



Gavesh Bhardwaj (GB Sir) Exp.: 17 yrs

Nitin Vijay (NV Sir) Managing Director Exp. : 18 yrs





Ram Ratan Dwivedi (RRD Sir) Joint Director Exp.: 20 yrs



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Durgesh Pandey (Pandey Sir) Sr. Faculty Exp.: 8 yrs



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ANSWER KEY

SECTION - A

- This section contains SIX (06) questions.
- Each question has FOUR options (A), (B), (C) and (D). ONE OR MORE THAN ONE of these four option(s) is (are) correct answer(s).
- For each question, choose the option(s) corresponding to (all) the correct answer(s).
- Answer to each question will be evaluated according to the following marking scheme:
 - Full Marks : +4 If only (all) the correct option(s) is(are) chosen;
 - Partial Marks : +3 If all the four options are correct but ONLY three options are chosen;
 - Partial Marks : +2 If three or more options are correct but ONLY two options are chosen, both of which are correct;
 - Partial Marks : +1 If two or more options are correct but ONLY one option is chosen and it is a correct option;

Zero Marks : 0 If unanswered;

Negative Marks : -2 In all other cases.

- For example, in a question, if (A), (B) and (D) are the ONLY three options corresponding to correct answers, then
 - choosing ONLY (A), (B) and (D) will get +4 marks;
 - choosing ONLY (A) and (B) will get +2 marks;
 - choosing ONLY (A) and (D) will get +2marks;
 - choosing ONLY (B) and (D) will get +2 marks;
 - choosing ONLY (A) will get +1 mark;

choosing ONLY (B) will get +1 mark;

choosing ONLY (D) will get +1 mark;

choosing no option(s) (i.e. the question is unanswered) will get 0 marks and choosing any other option(s) will get -2 marks.

SECTION - 1

1. The reaction sequence(s) that would lead to *o*-xylene as the major product is(are)



ANSWER KEY

Ans. AB



2. Correct option(s) for the following sequence of reactions is(are)



Ans. CD

Sol.



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3. For the following reaction $2X + Y \xrightarrow{k} P$

the rate of reaction is $\frac{d[P]}{dt} k[X]$. Two moles of **X** are mixed with one mole of **Y** to make 1.0 L of solution. At 50 s, 0.5 mole of **Y** is left in the reaction mixture. The correct statement(s) about the reaction is(are) (Use: ln 2 = 0.693) (A) The rate constant, k, of the reaction is $13.86 \times 10^{-4} s^{-1}$

(B) Half-life of X is 50 s.

(C) At 50s,
$$-\frac{d[X]}{dt} = 13.86 \times 10^{-3} \text{ mol } \text{L}^{-1} \text{ s}^{-1}$$

(D) At 100s, $-\frac{d[Y]}{dt} = 3.46 \times 10^{-3} \text{ mol } \text{L}^{-1} \text{ s}^{-1}$.

Ans. BCD Sol.

 $2X + Y \xrightarrow{k} P \qquad -\frac{1}{2} \frac{dX}{dt} = \frac{-dY}{dt} = \frac{dP}{dt} = k[X]$ $2 \qquad 1$ $1 \qquad 0.5 \qquad \frac{-dX}{dt} = 2[X] = kd[X]$ $t_{\frac{1}{2}}(A) = 50s = \frac{0.693}{k_d} \Rightarrow k_d = \frac{0.693}{50}$ $kd = 1.386 \times 10^{-2}$ $\frac{dX}{dt} = 1.386 \times 10^{-2} \times [X] \qquad [X] = \frac{1}{1} = 1$ $= 1.386 \times 10^{-2}$ $at 100 \text{ sec} \Rightarrow t_{\frac{3}{4}} = 100 \text{ sec}$ $[X] = \frac{A0}{2^2} = \frac{2}{4} = 0.5$ $\frac{dY}{dt} = k \times [X] = \frac{1.386}{2} \times 0.5 = 3.46 \times 10^{-2}$

4. Some standard electrode potentials at 298 K are given below:

Pb ²⁺ / Pb	-0.13V
Ni ²⁺ / Ni	-0.24V
Cd^{2_+} / Cd	-0.40V
Fe ²⁺ / Fe	_0 44 V

To a solution containing 0.001 M of X^{2+} and 0.1 M of Y^{2+} , the metal rods X and Y are inserted (at 298 K) and connected by a conducting wire. This resulted in dissolution of X. The correct combination(s) of X and Y, respectively, is (are)

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ANSWER KEY

Ans. ABC

5.

- **Sol.** $Pb^{+2} + 2e^{-} \longrightarrow Pb(s)$ $\begin{bmatrix} Ni^{+2} + 2e^{-} \longrightarrow Ni(s) & E^{\circ} = -0.13v \\ Cd^{2+} + 2e^{-} \longrightarrow Cd(s) & E^{\circ} = -0.24V \\ Fe^{+2} + 2e^{-} \longrightarrow Fe(s) & E^{\circ} = -0.4V \\ Fe^{+2} + 2e^{-} \longrightarrow Fe(s) & E^{\circ} = -0.44v \\ E = E^{\circ} - \frac{0.0591}{2} \log Q \\ -0.4 = -0.24 - \frac{0.0591}{2} \log Q \\ \end{bmatrix}$
 - The pair (s) of complexes wherein both exhibit tetrahedral geometry is (are) (Note: py = pyridine Given: Atomic numbers of Fe, Co, Ni and Cu are 26, 27, 28 and 29, respectively)
 - (A) $[FeCl_4]^-$ and $[Fe(CO)_4]^{2-}$ (B) $[CO(CO)_4]^-$ and $[COCl_4]^{2-}$
 - (C) $[Ni(CO)_{4}]$ and $[Ni(CN)_{4}]^{2-}$

(B) $\lfloor CO(CO)_4 \rfloor$ and $\lfloor COCl_4 \rfloor^2$ (D) $\lceil Cu(py)_4 \rceil^+$ and $\lceil Cu(CN)_4 \rceil^{3-1}$

- **A**, **B**, **D**
- Ans. A, B, D (1) $(FeCl_4)^- \longrightarrow$



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ANSWER KEY











- **6.** The correct statement(s) related to oxoacids of phosphorous is (are)
 - (A) Upon heating H_3PO_3 undergoes disproportionation reaction to produce H_3PO_4 and PH_3 .
 - (B) While H_3PO_3 can act as reducing agent, H_3PO_4 cannot.
 - (C) H_3PO_3 is a monobasic acid.
 - (D) The H atom of P–H bond in H_3PO_3 is not ionizable in water.



Ans. A, B, D

- (A) $H_3PO_3 \xrightarrow{disproportionation} H_3PO_4 + PH_3$
- (B) H_3PO_3 is reducing agent due to presence of P H bond
- (C) H_3PO_3 is dibasic acid due to presence of two –OH group
- (D) The H-atom of P–H bond is not Ionizable

Section – 2

ANSWER KEY

- This section contains THREE (03) question stems.
- There are TWO (02) questions corresponding to each question stem.
- The answer to each question is a NUMERICAL VALUE.
- For each question, enter the correct numerical value corresponding to the answer in the designated place using the mouse and the on-screen virtual numeric keypad.
- If the numerical value has more than two decimal places, truncate/round-off the value to TWO decimal places.
- Answer to each question will be evaluated according to the following marking scheme:
 Full Marks : +2 If ONLY the correct numerical value is entered at the designated place;
 Zero Marks : 0 In all other cases.

Question stem for Question Nos. 7 and 8

Question Stem

At 298 K, the limiting molar conductivity of a weak monobasic acid is 4×10^2 S cm² mol⁻¹. At 298 K, for an aqueous solution of the acid the degree of dissociation is α and the molar conductivity is $\mathbf{y} \times 10^2$ S cm² mol⁻¹. At 298 K, upon 20 times dilution with water, the molar conductivity of the solution becomes $3\mathbf{y} \times 10^2$ S cm² mol⁻¹.

7. The value of α is_____.

Ans. 0.22

8. The Value of y is _____.

Ans. 0.863

Sol. Λ^0_m (HA) = 4 × 10² Scm² / mol

 Λ^{C} HA = y × 10² Scm² / mol

$$\alpha = \frac{\Lambda^{\rm C}}{\Lambda^{\rm 0}}$$

When solution is diluted 20 times with water $\alpha_2 = 3\alpha_1$

$$k_{a} = \frac{C\alpha^{2}}{1-\alpha} = \frac{C}{20} \times \frac{(3\alpha)^{2}}{1-3\alpha}$$
$$\frac{1}{1-\alpha} = \frac{1}{20} \times \frac{9}{1-3\alpha}$$
$$\frac{1}{20-60\alpha} = 9-9\alpha$$
$$11 = (60-9)\alpha$$
$$\frac{11}{51} = \alpha = 0.22$$
$$\alpha = 0.22$$
$$\alpha = \frac{\Lambda^{c}}{\Lambda^{0}} = \frac{Y \times 10^{2}}{4 \times 10^{2}} = \frac{11}{51}$$
$$y = \frac{44}{51} = 0.863$$

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Question stem for Question Nos. 9 and 10

ANSWER KEY

Question Stem

Reaction of \mathbf{x} g of Sn with HCl quantitatively produced a salt. Entire amount of the salt reacted with \mathbf{y} g of nitrobenzene in the presence of required amount of HCl to produce 1.29 g of an organic salt (quantitatively).

(Use Molar masses (in g mol⁻¹) of H, C, N, O, Cl and Sn as 1, 12, 14, 16, 35 and 119, respectively).

9. The value of x is_____.

Ans. 3.57

- **10.** The value of y is _____.
- Ans. 1.23
- Sol. $Sn + 2HCI \longrightarrow SnCl_2 \longrightarrow x / 119mol$ NO₂ Sn/HCl



Question stem for Question Nos. 11 and 12

Question Stem

A sample (5.6 g) containing iron is completely dissolved in cold dilute HCl to prepare a 250 mL of solution. Titration of 25.0 mL of this solution requires 12.5 mLof 0.03 M KMnO₄ solution to reach the end point. Number of moles of Fe²⁺ present in 250 mL solution is $\mathbf{x} \times 10^{-2}$ (consider complete dissolution of FeCl₂). The amount of iron present in the sample is y% by weight.

(Assume: $KMnO_4$ reacts only with Fe^{2+} in the solution Use: Molar mass of iron as 56 g mol⁻¹)

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ANSWER KEY

11. The value of x is _____.

Ans. 1.875

12. The value of y is _____.

Ans. 18.75

Sol. Fe + 2HCl \longrightarrow FeCl₂ + H₂ meq of KMnO₄ = meq of Fe⁺² 12.5 × 0.03 × 5 = 1.875 = meq of Fe⁺¹ in 25ml meq of Fe⁺² = m.moles of Fe⁺¹ in 250 ml = 18.75 moles = = 18.75 × 10⁻³ = x × 10⁻² wt = 18.75 × 10⁻³ × 56 = 1.05 gm x = 1.875 % of Fe⁺² = $\frac{1.05}{5.6}$ × 100 = 18.75

Section – 3

- This section contains TWO (02) paragraphs. Based on each paragraph, there are TWO (02) questions.
- Each question has FOUR options (A), (B), (C) and (D). ONLY ONE of these four options is the correct answer.
- For each question, choose the option corresponding to the correct answer.
- Answer to each question will be evaluated according to the following marking scheme:

Full Marks : +3 If ONLY the correct option is chosen;

Zero Marks : 0 If none of the options is chosen (i.e. the question is unanswered); Negative Marks : -1 In all other cases.

Paragraph

The amount of enrgy required to break a bond is same as the amount energy released when the same bond is formed. In gaseous state, the energy required for homolytic cleavage of a bond is called Bond Dissociation Energy (BDE) or Bond Strength. BDE is affected by s-character of the bond and the stability of the radicals formd. Shorter bonds are typically stronger bonds. BDEs for some bonds are given below:

$$H_{3}C - H(g) \longrightarrow H_{3}C^{\bullet}(g) + H^{\bullet}(g) \quad \Delta H^{\circ} = 105 \text{ kcal mol}^{-1}$$

$$CI - CI(g) \longrightarrow CI^{\bullet}(g) + CI^{\bullet}(g) \quad \Delta H^{\circ} = 58 \text{ kcal mol}^{-1}$$

$$H_3C - CI(g) \longrightarrow H_3C^{\bullet}(g) + CI^{\bullet}(g) \Delta H^{\circ} = 85 \text{ kcal mol}^{-1}$$

$$H - Cl(g) \longrightarrow H(g) + Cl(g) \Delta H^{\circ} = 103 \text{ kcal mol}^{-1}$$

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13. Correct match of the **C**-**H** bonds (shown in bold in Column **J** with their BDE in Column **K** is)

	Column J	Column K
	Molecule	BDE (kcal mol ⁻¹)
	(P) H -CH(CH ₃) ₂	(i) 132
	(Q) H –CH ₂ Ph	(ii) 110
	(R) \mathbf{H} -CH=CH ₂	(iii) 95
	(S) H −C ≡ CH	(iv) 88
	(A) P-iii, Q-iv, R-ii,	S-i
	(B) P-i, Q-ii, R-iii, S	-iv
	(C) P-iii, Q–ii, R–i, S	-iv
	(D) P-ii, Q–i, R–iv, S	iii
Ans.	D	
Sol.	$CH_3 - H \longrightarrow CH_3 +$	$\Delta H^{\circ} = 105 \text{ K Cal}$
	5	(A)
	1	
	B.D.E α stability of fr	reeradical
	$CH_{4} \longrightarrow CH_{2} + H$	$\Delta H^{\circ} = 105$
	$HCI \longrightarrow H+CI \qquad \Lambda$	H° = 103
	•	
	$CI + CH_4 \longrightarrow HCI +$	$CH_3 \qquad \Delta H^\circ = 105 - 103$
		= 2
		endothermic
	$\dot{C}H_3 + Cl_2 \longrightarrow CH_3$	CI + ĊI
	$\dot{C}H_3 + \dot{C}I \longrightarrow CH_3C$	$\Delta H^{\circ} = -85$
	Cl−Cl→2Ċl	$\Delta H^{o} = 58$
		<u> </u>
	$\dot{C}H_3 + Cl_2 \longrightarrow CH_3Q$	$CI + CI \qquad \Delta H^{\circ} = -27$
	a	exothermic
	final reaction	
	$CH_4 + Cl_2 \longrightarrow CH_3C$	CI + HCI
	$\Delta H^{o} = 105 + 5$	8-85-103
	= -25	
	CH ₂ > •	H_2 > $H_2 = CH_2 > C = CH_2$
		ĊH₃
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14. For the following reaction

 $CH_4(g) + CI_2(g) \xrightarrow{light} CH_3CI(g) + HCI(g)$

the correct statement is

- (A) Initiation step is exothermic with $\Delta H^{\circ} = -58 \text{ kcal mol}^{-1}$
- (B) Propagation step involving ${}^{\circ}CH_{3}$ formation is exothermic with $\Delta H^{\circ} = -2 \text{ kcal mol}^{-1}$
- (C) Propagation step involving CH_3CI formation is endothermic with $\Delta H^{\circ} = +27 \text{ kcal mol}^{-1}$
- (D) the recaiton is exothermic with $\Delta H^{\circ} = -25 \text{ kcal mol}^{-1}$

Ans. D

Sol. $H_3C - H \rightarrow H_3\dot{C} + \dot{H}_2 \Delta H^\circ = 105 \text{ kcal/mole}$

 $CI - CI \rightarrow CI + CI_{,} \Delta H^{\circ} = 58 \text{ kcal/mole}$ $H_{3}C - CI \rightarrow H_{3}C + CI_{,} \Delta H^{\circ} = 85 \text{ kcal/mole}$ $H - CI \rightarrow \dot{H} + \dot{C}I_{,} \Delta H^{\circ} = 103 \text{ kcal/mole}$ $CH_{4} + CI_{2} \rightarrow CH_{3} - CI + HCI$ **Step-I** Chain intiationstep

 $CI - CI \rightarrow 2CI$ (+58)

Step-II Chain propogation step

$$CH_3 - H + \dot{CI} \rightarrow H - CI + \dot{CH}_3$$

$$\vdots$$

 $CH_3 + CI - CI \rightarrow CH_3 - CI + CI$

Total energy gain = $105 + 58 = 163 \text{ kcal mol}^{-1}$ Total energy released = $103 + 85 = 188 \text{ kcal mol}^{-1}$

Paragraph

The reaction of $K_3[Fe(CN)_6]$ with freshly prepared $FeSO_4$ solution produces a dark blue precipitate called Turnbull's blue. Reaction of $K_4[Fe(CN)_6]$ with the $FeSO_4$ soluton in complete absence of air produces a white preipitate **X**, which turns blue in air. Mixing the $FeSO_4$ solution with NaNO₃, followed by a slow addition of concentrated H_2SO_4 through the side of the test tube produces a brown ring.

15. Precipitate **X** is

(A) $Fe_4[Fe(CN)_6]_3$ (B) $Fe[Fe(CN)_6]$ (C) $K_2Fe[Fe(CN)_6]$ (D) $KFe[Fe(CN)_6]$

Ans. C

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(B) $[Fe(NO)_{2}(H_{2}O)_{4}]^{3+}$ (A) $[Fe(NO)_{2}(SO_{4})_{2}]^{2-}$

- (D) $[Fe(NO)(H_2O)_5]^{2+}$ (C) $[Fe(NO)_4(SO_4)_2]$
- D Ans.

(3)

(1) K_3 [Fe(CN)₆] + FeSO₄ \longrightarrow Fe₃[Fe(CN)₂]₂

- Turnbull's blue
- (2) $K_4 [Fe(CN)_6] + FeSO_4 \longrightarrow K_3Fe[Fe(CN)_6]$





 $Fe_4[Fe(CN)_6]_3$ Or K Fe [Fe(CN)₆] Prussian blue $FeSO_4 + NaNO_3 + H_2SO_4 \longrightarrow [Fe(H_2O)_5NO]^{+2}$ Conc.

SECTION 4

- This section contains THREE (03) questions.
- The answer to each question is a NON-NEGATIVE INTEGER.
- For each question, enter the correct integer corresponding to the answer using the mouse and the on-screen virtual numeric keypad in the place designated to enter the answer.
- Answer to each question will be evaluated according to the following marking scheme: Full Marks : +4 If ONLY the correct integer is entered; : 0 In all other cases. Zero Marks
- 17. One mole of an ideal gas at 900 K, undergoes two reversible processes, I followed by II, as shown below. If the work done by the gas in the two processes are same, the value of In $\frac{V_3}{V}$ is

 $2250 - - - - (p_1, V_1)$ $U_{R}(K)$ $450 - - (p_2, V_2) - (p_3, V_3)$ $C(1)(c^1 - c^1) + (c^1 - c^1)$

(U: internal energy, S: entropy, p:pressure, V: volume, R:gas constant) (Given: molar heat capacity at constant volume, $C_{v,m}$ of the gas is $\frac{5}{2}R$)

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ANSWER KEY

Ans. 10



$$\begin{split} & \text{In a reversible adiabatic exp.} \\ & \frac{\Delta U}{R}(I) = 450 - 2250 = -1800 \\ & \Delta U = -1800 \text{R} \\ & \text{Process (II) is an Isothermal exp.} \\ & W_{(I)} = W_{(II)} \\ & nC_v(T_2 - T_1) = -nRT_2 ln \frac{V_3}{V_2} \\ & C_v(T_2 - T_1) = -RT_2 ln \frac{V_3}{V_2} \\ & W_{(I)} = \Delta U = -1800 \text{R} \\ & nC_v(T_2 - T_1) = -1800 \text{R} \\ & 1 \times \frac{5}{3} \text{R}(T_2 - 900) = -1800 \text{R} \\ & 1 \times \frac{5}{3} \text{R}(T_2 - 900) = -1800 \text{R} \\ & \frac{-3600}{5} = T_2 - 900 \\ & T_2 = 180 \text{K} \\ & -1800 \text{R} = -RT_2 ln \frac{V_3}{V_2} \\ & -1800 \text{R} = -180 ln \frac{V_3}{V_2} \end{split}$$

$$ln\frac{V_3}{V_2} = 10$$

18. Consider a helium (He) atom that absorbs a photon of wavelength 330 nm. The change in the velocity (in cm s⁻¹) of He atom after the photon absorption is _____. (Assume: Momentum is conserved when photon is absorbed.) Use: Planck consant = 6.6×10^{-34} Js, Avogadro number = 6×10^{23} mol⁻¹, Molar mass of He = 4gmol⁻¹)





Ans. 30

Sol. $\lambda = \frac{h}{m(\Delta v)}$ $\Delta v = \frac{h}{m\lambda} = \frac{6.6 \times 10^{-34} N_A}{4 \times 10^{-3} \times 330 \times 10^{-9}}$ = 0.3 m/sec = 30 Cm/sec

19. Ozonolysis of ClO₂ produces an oxide of chlorine. The average oxidation state of chlorine in this oxide is__.

ANSWER KEY

- Ans. 6
- **Sol.** CIO_2 contains an odd electron and is paramagnetic. It reacts with ozone to give O_2 and CI_2O_6 . $2CIO_2 + 2O_3 \longrightarrow CI_2O_6 + 2O_2$

In Cl_2O_6 , the average oxidation state of Cl is +6.



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